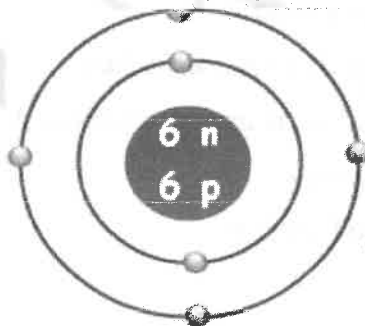


DELHI PUBLIC SCHOOL, DURGAPUR
QUESTION BANK FOR PERIODIC ASSESSMENT -II (2018-19)
CLASS-VIII

SUBJECT: CHEMISTRY

TOPIC : ATOMIC STRUCTURE

- Q1. Write the main postulates of Dalton's atomic theory.
- Q2. Write about the contradiction of Dalton's atomic theory by the Modern atomic theory.
- Q3. Name the scientist who discovered – i. Cathode rays ii. Electrons
 iii. Protons iv. Atomic nucleus v. Neutrons
- Q4. State the atomic number and mass number of the following:
i. Lithium ii. Neon iii. Silicon iv. Potassium v. Chlorine
- Q5. Differentiate between:
i. Relative atomic mass and Relative molecular mass
ii. Isotope and Isobar
iii. Stable and unstable electronic configuration
iv. Electrovalency and covalency
- Q6. Write the electronic configuration of :
i. Boron ii. Fluorine iii. Magnesium iv. Phosphorus v. Argon
- Q7. What type of compounds are the following -i) CaO ii) $MgCl_2$ iii) C_2H_4 iv) H_2O
- Q8. Which isotope of hydrogen has two neutrons in its nucleus? Name the isotope.
- Q9. Name the isotope of hydrogen which is i) radio active in nature, ii) used to prepare heavy water. How many electrons does each of them have?
- Q10. Calculate the maximum capacity of electrons in M shell.
- Q11. The diagram represents one atom of carbon .Answer the following questions



- a. Number of valence electrons
b. Atomic number of carbon

c. Electronic configuration of carbon

Q12. Both helium and beryllium have two valence electrons. Helium is a noble gas but beryllium is not. Give reason.

Q13. An element has 3 electrons in the M shell. What is its atomic number?

Q14. element X has 17 electrons and element Y has 16 electrons. Which one is more reactive and why?

TOPIC : LANGUAGE OF CHEMISTRY

Q1. Write the formula of the following compound

- | | | |
|------------------------|-------------------------|------------------------|
| i. Potassium chloride | ii. Potassium hydroxide | iii. Sodium nitrate |
| iv. Sodium bicarbonate | v. Calcium bisulphate | vi. Calcium nitride |
| vii. Magnesium oxide | viii. Zinc hydroxide | ix. Copper(I) sulphide |
| x. Iron(II) chloride | | |

Q2. Balance the following equations

- Sodium + Water \rightarrow Sodium hydroxide + Hydrogen
- Sodium hydroxide + Sulphuric acid \rightarrow Sodium sulphate + Water
- Calcium bicarbonate + Carbon dioxide \rightarrow Calcium carbonate + Water
- Zinc + Sodium hydroxide \rightarrow Sodium zincate + Hydrogen
- Lead(II) nitrate \rightarrow Lead(II) oxide + Nitrogen oxide + Oxygen
- Aluminium oxide + Hydrochloric acid \rightarrow Aluminium chloride + Water

TOPIC : COMBUSTION AND FLAME

Q1. Assign reasons:

- Camphor burns with a flame but a piece of charcoal glows
- Fire produced by burning oil or petrol cannot be extinguished by water.
- White Phosphorus bursts into flame during summer season.
- Hydrogen has a high calorific value, yet it is not used as a domestic fuel.
- Kerosene burns with a blue flame in a wick stove whereas it imparts a yellow flame in lamp.
- Petrol is considered as a better fuel than coal or charcoal.
- Fire produced by electricity cannot be extinguished by water.

Q3. State four main characteristics of a good fuel.

Q4. Describe the mechanism of Soda Acid fire Extinguisher.

Q5. Draw a neat and labelled diagram of different zones of candle.

Q6. In which zone of candle flame, unburnt carbon particles are found. Name the hottest zone of the candle flame.

TOPIC : COAL AND PETROLEUM

Q1. Describe how coal was formed? What is the process called?

Q2. What is destructive distillation of coal? What are its products?

Q3. State use of coke.

Q4. What are the constituents of coal gas? State one use of coal gas.

Q5. Which is the best variety of coal and why?

SYLLABUS

UNIT 1,2 & 3 : ATOMIC STRUCTURE

UNIT 4 & 5 : LANGUAGE OF CHEMISTRY

UNIT 6 & 7 : COMBUSTION AND FLAME

UNIT 7 : COAL AND PETROLEUM