

CHEMISTRY

MONTH	CHAPTER NO.	CHAPTER	CONTENTS
April	I	Some Basic Concepts of Chemistry	<ul style="list-style-type: none"> laws of chemical combination, Dalton's atomic theory: concept of elements, atoms and molecules. Atomic and molecular masses mole concept and molar mass percentage composition, empirical and molecular formula chemical reactions, stoichiometry and calculations based on stoichiometry.
May-June	II	Structure of Atom	<ul style="list-style-type: none"> Bohr's model and its limitations. concept of shells and subshells dual nature of matter and light, de Broglie's relationship Heisenberg uncertainty principle concept of orbitals, quantum numbers, shapes of s, p and d orbitals rules for filling electrons in orbitals - Aufbau principle, Pauli's exclusion principle and Hund's rule electronic configuration of atoms, stability of half-filled and completely filled orbitals.
July	UT I		
July	III	Classification of Elements and Periodicity in Properties	<ul style="list-style-type: none"> Modern periodic law and the present form of periodic table periodic trends in properties of elements -atomic radii, ionic radii, inert gas radii, Ionization enthalpy, electron gain enthalpy, electronegativity, valency. Nomenclature of elements with atomic number greater than 100
July	IV	Chemical Bonding and Molecular structure	<ul style="list-style-type: none"> Valence electrons, ionic bond, covalent bond, bond parameters, Lewis structure polar character of covalent bond, covalent character of ionic bond valence bond theory, resonance geometry of covalent molecules, VSEPR theory concept of hybridization, involving s, p and d orbitals and shapes of some simple molecules molecular orbital theory of homonuclear diatomic molecules(qualitative idea only), hydrogen bond.
July	V	States of Matter: Gases, Liquids.	<p>Three states of matter, intermolecular interactions, types of bonding, melting and boiling points, role of gas laws in elucidating the concept of the molecule</p> <ul style="list-style-type: none"> Boyle's law, Charles law, Gay Lussac's law, Avogadro's law, ideal behaviour, empirical derivation of gas equation, Avogadro's number, ideal gas equation. Deviation from ideal behaviour liquefaction of gases, critical temperature, kinetic energy and molecular speeds (elementary idea) Liquid State: vapour pressure, viscosity and surface tension (qualitative idea only, no mathematical derivations)
August	VI	Thermodynamics	<ul style="list-style-type: none"> Concepts of System and types of systems, surroundings, work, heat, energy, extensive and intensive properties, state functions. First law of thermodynamics -internal energy and enthalpy, heat capacity and specific heat, measurement of ΔU and ΔH.

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August	VI	Thermodynamics	<ul style="list-style-type: none"> Hess's law of constant heat summation, enthalpy of bond dissociation, combustion, formation, atomization, sublimation, phase transition, ionization, solution and dilution. Second law of Thermodynamics (brief introduction). Introduction of entropy as a state function, Gibb's energy change for spontaneous and non- spontaneous processes, criteria for equilibrium. Third law of thermodynamics (brief introduction).
August	XII	Organic Chemistry - Some Basic Principles and Techniques	<ul style="list-style-type: none"> Classification and IUPAC nomenclature of organic compounds. Isomerism
September	REVISION AND BLOCK TEST 1		
October	XII	Organic Chemistry - Some Basic Principles and Techniques (CONTINUED)	<ul style="list-style-type: none"> Electronic displacements in a covalent bond: inductive effect, electromeric effect, resonance and hyper conjugation. Homolytic and heterolytic fission of a covalent bond: free radicals, carbocations, carbanions, electrophiles and nucleophiles, types of organic reactions.
November	IX	Hydrogen	<ul style="list-style-type: none"> Position of hydrogen in periodic table, occurrence, isotopes, preparation, properties and uses of hydrogen, hydrides-ionic covalent and interstitial; physical and chemical properties of water, heavy water, hydrogen peroxide -preparation, reactions and structure and use; hydrogen as a fuel.
	UT 2		
	XIII	Hydrocarbons	<ul style="list-style-type: none"> Alkanes - Nomenclature, isomerism, conformation (ethane only), physical properties, chemical reactions including free radical mechanism of halogenation, combustion and pyrolysis. Alkenes - Nomenclature, structure of double bond (ethene), geometrical isomerism, physical properties, methods of preparation chemical reactions: addition of hydrogen, halogen, water, hydrogen halides (Markownikov's addition and peroxide effect), ozonolysis, oxidation, mechanism of electrophilic addition. Alkynes - Nomenclature, structure of triple bond (ethyne), physical properties, methods of preparation, chemical reactions: acidic character of alkynes, addition reaction of - hydrogen, halogens, hydrogen halides and water. Aromatic Hydrocarbons: Introduction, IUPAC nomenclature, benzene: resonance, aromaticity, chemical properties: mechanism of electrophilic substitution. Nitration, sulphonation, halogenation, Friedel Craft's alkylation and acylation, directive influence of functional group in monosubstituted benzene. Carcinogenicity and toxicity.
December	VII	Equilibrium	<ul style="list-style-type: none"> Equilibrium in physical and chemical processes, dynamic nature of equilibrium law of mass action, equilibrium constant factors affecting equilibrium- Le Chatelier's principle,

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December			<ul style="list-style-type: none">ionic equilibrium- ionization of acids and bases, strong and weak electrolytes, degree of ionization, ionization of poly basic acids, acid strength, concept of pHHenderson Equation, hydrolysis of salts (elementary idea), buffer solution, solubility product, common ion effect (with illustrative examples).
	VIII	Redox Reactions	<ul style="list-style-type: none">Concept of oxidation and reduction, redox reactions, oxidation number,balancing redox reactions, in terms of loss and gain of electrons and change in oxidation number.Electrochemical cellsapplications of redox reactions.
January	UT 3		
	X	s - block Elements (Alkali and Alkaline Earth Metals)	<p>Group 1 and Group 2 Elements :-General introduction, electronic configuration, occurrence, anomalous properties of the first element of each group, diagonal relationship,</p> <ul style="list-style-type: none">trends in the variation of properties (such as ionization enthalpy, atomic and ionic radii), trends in chemical reactivity with oxygen, water, hydrogen and halogens, uses.Preparation and Properties of Some Important Compounds: Sodium Carbonate, Sodium Chloride, Sodium Hydroxide and Sodium Hydrogencarbonate,Biological importance of Sodium and Potassium. Calcium Oxide and Calcium Carbonate and their industrial uses, biological importance of Magnesium and Calcium.
January	XI	p-block Elements	<ul style="list-style-type: none">General Introduction to p -Block Elements Group 13 Elements: General introduction, electronic configuration, occurrence, variation of properties, oxidation states, trends in chemical reactivity, anomalous properties of first element of the group,Boron - physical and chemical properties, some important compounds, Borax, Boric acid, Boron Hydrides, Aluminium: Reactions with acids and alkalis, usesGroup 14 Elements: General introduction, electronic configuration, occurrence, variation of properties, oxidation states, trends in chemical reactivity, anomalous behaviour of first elements.Carbon-catenation, allotropic forms, physical and chemical properties; uses of some important compounds: oxides. Important compounds of Silicon and a few uses: Silicon Tetrachloride, Silicones, Silicates and Zeolites, their uses.
	XIV	Environmental Chemistry	<ul style="list-style-type: none">Environmental pollution - air, water and soil pollution,chemical reactions in atmosphere, smog, major atmospheric pollutants, acid rainozone and its reactions, effects of depletion of ozone layer, greenhouse effect and global warming- pollution due to industrial wastes,green chemistry as an alternative tool for reducing pollution, strategies for control of environmental pollution concentration of either of the ions.
	Revision		
	Revision and BT-2		
February			
Theory Paper 70 Marks + Practical Work 30 Marks			



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Practical work: Project File and Viva

Practical Marks: 30 Marks

Exam	Topics
Block Test 1	1. To Identify the acidic radical in the given salt sample.
	(i) Carbonate, Sulphide, Nitrite
	(ii) Chloride, Bromide, Iodide, Nitrate, Acetate
	(iii) Sulphate
	2. To Identify the basic radical in the given salt sample.
	(i) Group Zero radical: Ammonium
	(ii) Group I radicals: Lead,
	(iii) Group II radicals: copper
Block Test 2	3. To identify the basic radical (continued)
	(i) Group III radicals: Aluminium, Iron(III)
	(ii) Group IV radicals: Cobalt, Nickel, Manganese, Zinc
	(iii) Group V radicals: Barium, Calcium, Strontium
	4. TITRATION
	To calculate the strength of NaOH by titrating it against standard oxalic acid
	To calculate the strength of HCl by titrating it against standard sodium carbonate.
	5. INVESTIGATORY PROJECT
	Any one topic either from lab manual or any other relevant topic based on CBSE syllabus.

Assessment Criteria

Heading	Marks
Identification of acid and basic radical in the given salt sample	10
Titration	10
Project	5
Lab file + Viva	5
Total Marks	30